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Raoult's Law - How

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Nonvolatile Solute 2.3

Vapor Pressure, IMFs,

and Boiling Point

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Vapor Pressure

Partial Pressures

/u0026 Vapor

Pressure: Crash

Course Chemistry

#15 ~~Vapor pressure |~~

~~States of matter and~~

~~intermolecular forces~~

~~| Chemistry | Khan~~

Academy Vapor

Pressure 101 Vapor

Pressure lowering :

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Law Vapor Pressure

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Basic Introduction,

Normal Boiling Point,

$\Delta H_{vap}$  Clausius

Clapeyron Equation -

Chemistry 13.3 Vapor

Pressure Depression

and Raoult's Law

~~Vapor Pressure,~~

~~Equilibrium Vapor~~

~~Pressure, and Relative~~

~~Humidity~~ Vapor

Pressure Relative

Humidity vs

Dewpoint Raoult's

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Law With Example

Problem How does a  
Pressure Cooker

Work? - Science for

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Vapor Pressure and

Boiling Boiling Point

Raoults Law and

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Vapour Pressure -

The States Of Matter

(Part 25) Vapor

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Pressure of Liquids

Vapor Pressure and

Boiling Boiling;

Atmospheric

Pressure, and Vapor

Pressure Vapor

Pressure Equilibrium

Demonstration of

Vapor Pressure

Humidity 102: Vapor



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~~Vapor Pressure~~

~~VPD Chart FSc~~

Chemistry Book1, CH

9, LEC 10: Vapor

Pressure of Ideal

Solutions 11.5 Vapor

Pressure What Is The

Vapor Pressure

Vapor pressure (or

vapour pressure in

British English; see

spelling differences)

or equilibrium vapor

pressure is defined as

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vapor pressure exerted by a vapor in thermodynamic equilibrium with its condensed phases (solid or liquid) at a given temperature in a closed system. The equilibrium vapor pressure is an indication of a liquid's evaporation rate. It relates to the tendency of particles

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Vapor pressure -

Wikipedia

The vapor pressure of a liquid is the equilibrium pressure of a vapor above its liquid (or solid); that is, the pressure of the vapor resulting from evaporation of a liquid (or solid) above a sample of the liquid

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What Is The

(or solid) in a closed  
container.

Of Solvent In An

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Vapor Pressure -

Purdue Chemistry

Vapour pressure,  
pressure exerted by a  
vapour when the  
vapour is in  
equilibrium with the  
liquid or solid form,  
or both, of the same  
substance—i.e., when  
conditions are such

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vapour pressure

of a substance  
can exist in both or in  
all three phases.

Vapour pressure is a  
measure of the

tendency of a  
material to change  
into the gaseous or  
vapour state, and it  
increases with  
temperature.

vapour pressure |

Definition & Facts |

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Britannica

Vapor pressure is the pressure exerted by a vapor which is in

thermodynamic

equilibrium with its

condensed phases

(solid or liquid) in a

closed system at a

given temperature.

The equilibrium - in

other words, steady

state - between

evaporation and

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What Is The

condensation occurs  
when:

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Vapor Pressure of

Water. Calculator |

Definition | Formulas

A The vapor pressure  
curve of water

intersects the  $P =$

1000 mmHg line at

about  $110^{\circ}\text{C}$ ; this is

therefore the boiling

point of water at

1000 mmHg. B The

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What Is The

Vapor Pressure

corresponding to  
250 ° C intersects the

vapor pressure curve

of mercury at P = 75

mmHg. Hence this is

the pressure required

for mercury to boil at

250 ° C.

11.5: Vapor Pressure

- Chemistry

LibreTexts

Vapor pressure is the



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What Is The

pressure exerted by the vapor molecules of a substance in a closed system. It occurs at equilibrium, i.e., when the molecules are both vaporizing and condensing at the same rate at a particular pressure.

Vapor Pressure

Calculator | Clausius-

*Page 17/34*

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What Is The

Clapeyron Equation

Vapour pressure can be defined as

pressure formed by

the vapor of the

liquid (or solid) over

the surface of the

liquid. This pressure

is formed in a

thermodynamic

equilibrium state in a

closed container at a

certain temperature.

Liquid ' s evaporation

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What Is The

rate is identified by  
the equilibrium vapor  
pressure.

Aqueous

Vapour Pressure -

Definition, Raoult's  
Law, Formula, Videos

The boiling point is  
defined as the  
temperature at which  
the saturated vapor  
pressure of a liquid is  
equal to the  
surrounding

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vapor pressure of a liquid at atmospheric pressure.

For water, the vapor pressure reaches the standard sea level atmospheric pressure of 760 mmHg at 100 ° C.

Vapor Pressure -

Georgia State

University

Vapor pressure is constant when there is an equilibrium of

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water molecules moving between the liquid phase and the gaseous phase, in a closed container. The vapor pressure of a liquid is the point at which equilibrium pressure is reached, in a closed container, between molecules leaving the liquid and going into the gaseous phase and

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Of Solvent In An  
Aqueous  
Solution

molecules leaving the gaseous phase and entering the liquid phase.

Solution

Vapor Pressure and  
Water - [usgs.gov](http://usgs.gov)

Vapor pressure is the pressure exerted by vapor which is in equilibrium with its condensed form (liquid or solid phase).  
But when considering

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What Is The

vapor pressure,  
the system where the  
vapor exists should  
be a closed system  
with a constant  
temperature.

Difference Between  
Partial Pressure and  
Vapor Pressure ...

The boiling point of a  
substance is the  
temperature at which  
the vapor pressure of

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Of Solvent In An  
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the liquid equals the pressure surrounding the liquid, and the liquid changes into a vapor. The boiling point of a liquid varies depending upon the surrounding pressure. A liquid in a partial vacuum has a lower boiling point than when that liquid is at atmospheric pressure.



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## What Is The

### Vapor Pressure

#### Water Boiling Points

#### at Vacuum Pressure

Vapor pressure is the pressure of the vapor over a liquid (and some solids) at equilibrium. Now, what does that definition mean? I'm going to go through some explanation steps that, hopefully, give you a correct

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idea of vapor pressure. 1) Imagine a closed box of several liters in size.

Solution

ChemTeam: Vapor Pressure

By definition, vapor pressure is the amount of pressure within a vapor or gas when the substance is in an equilibrium state. In other words,

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When a liquid or solid is in a closed container and some molecules evaporate while others return to the liquid or solid state, the pressure that can be measured within that container relates to the vapor.

What Is Vapor  
Pressure? - wiseGEEK  
Vapor pressure is the

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vapor pressure caused by the evaporation of liquids. Three common factors that influence vapor pressure are surface area, intermolecular forces and temperature. The vapor pressure of a molecule differs at different temperatures. vapor pressure liquid evaporation

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Vapor Pressure  
Vapor Pressure -  
Concept - Chemistry  
Video by Brightstorm  
The pressure exerted

by a vapor on the  
solid or liquid phase  
with which it is in  
equilibrium. At  
pressures lower than  
the vapor pressure,  
atoms or molecules of  
the liquid or solid  
being vaporized can

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escape from the

surface of the liquid

or solid. At the vapor

pressure, they cannot

escape because the

two phases are in

equilibrium.

Vapor pressure -

definition of vapor

pressure by The Free

...

The vapor pressure is

a property of the

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vapor pressure of a substance and is constant at a given temperature. It increases when temperature increases. 2.) The boiling point of a substance is the temperature at which the vapor pressure of the liquid equals the pressure surrounding the liquid.

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What Is The

Vapor Pressure

(video) | States of  
matter | Khan

Academy

In chemistry, vapor pressure is the pressure that is exerted on the walls of a sealed container when a substance in it evaporates (converts to a gas). To find the vapor pressure at a given



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temperature, use the  
Clausius-Clapeyron  
equation:  $\ln(P_1/P_2) =$   
 $(-H_{\text{vap}}/R)((1/T_2) -$   
 $(1/T_1))$ .

3 Ways to Calculate  
Vapor Pressure -  
wikiHow

Definition of vapor  
pressure : the  
pressure exerted by a  
vapor that is in  
equilibrium with its

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What Is The

solid or liquid form —  
called also vapor  
tension Examples of  
vapor pressure in a  
Sentence

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