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answer choices Mole ratios are used as a conversion factor to convert from one substance to another. Mole ratios come from the coefficients in a balanced chemical equation.

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Stoichiometry Worksheet #1 Answers 1.

Given the following equation:  $2 \text{C}_4\text{H}_{10} + 13 \text{O}_2 \rightarrow 8 \text{CO}_2 + 10 \text{H}_2\text{O}$ ,  
show what the following molar ratios  
should be. a.  $\text{C}_4\text{H}_{10} : \text{CO}_2$ . [Filename:

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substances atomic mass in grams for mass.  
The % of an element in a compound by  
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make some sandwiches! \_ Introduction:

When we bake/cook something, we use a specific amount of each ingredient.

Imagine if you made a batch of cookies and used way too many eggs, or not enough sugar.

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Q. What is the percent yield if 0.856 g of  $\text{NH}_3$  is actually obtained in the lab during



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the following reaction:  $4\text{NH}_3 + 5\text{O}_2 \rightarrow 4\text{NO} + 6\text{H}_2\text{O}$  How many grams of NO are formed if 6.30g of ammonia react with 1.80g of oxygen?

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W/ answers Website Upload Unit 6:

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Answers and Stoichiometry (Link to  
Prentice Hall Text: Chapters 7, 8 & 9)  
Name: \_\_\_\_\_ Date Due Assignments Page  
Number: Problem Numbers Assignment 1:  
Interpreting Equations and Balancing 232:  
32, 39

## **Unit 6: Reactions and Stoichiometry**

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**Begin Unit 9:** Stoichiometry Intro Mole calculations summary: Assignments: Due 4/29 or 4/30: Stoichiometry \*work shown for #1 \* Stoichiometry key Due Tues and Wed: Unit 8 Pretest Test on 4/25 and 4/26 Tutorial Videos: mole-mole [HERE](#) mole-mass and mass-mole [HERE](#) mass-mass [HERE](#)

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long chemistry unit available by chapter.

Answer key is included. Chapter 8

Objectives: 1. Determine mole ratios from a balanced chemical equation. 2. Solve stoichiometry problems. 3. Identify the limiting reactant

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UNIT 9. STOICHIOMETRY The key to a



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**Answers**  
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