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combined gas law  
describes the relationship  
among the temperature,  
volume, and pressure of a  
gas when the number of  
particles is constant  
freezing point of water in  
Fahrenheit and Celcius  
32 degrees F, 0 degrees C

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## Section 32 Gas Laws Answers

~~chapter 3 section 3.2~~

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Section 3.2 The Gas

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pressure. the result of a force distributed over an area. which unit is used to express amounts of pressure. pascal and kilopascal. what causes the pressure in a closed container of gas?

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## Section 32 Gas

### collisions between

### particles of a gas and the

### walls of the container.

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### Laws Answers

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In 1662, Robert Boyle discovered the correlation between Pressure (P) and Volume (V) (assuming Temperature (T) and Amount of Gas (n) remain constant):  $P \propto \frac{1}{V}$  or  $PV = x$ . where x is a constant depending on amount of gas at a given

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temperature. Pressure is inversely proportional to Volume.

~~Gas Laws: Overview~~  
~~Chemistry LibreTexts~~  
Ideal Gas Law. The Ideal Gas Law mathematically relates the pressure, volume, amount and temperature of a gas with the equation:  $\text{pressure} \times \text{volume} = \text{moles} \times \text{ideal gas constant} \times$

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temperature;  $PV = nRT$ .

The Ideal Gas Law is ideal because it ignores interactions between the gas particles in order to simplify the equation.

~~Gas Laws (video lessons, examples and solutions)~~

This gas laws worksheet comprises Boyles law, Charles law and pressure law. It will help and challenge learners to

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understand how to solve problems involving gas laws.

### ~~GAS LAWS~~ ~~WORKSHEET WITH~~ ~~ANSWERS~~ | Teaching Resources

A.-C. Charles (1746 – 1823)—states that, at constant pressure, the volume  $V$  of a gas is directly proportional to its absolute (Kelvin)

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temperature  $T$ , or  $V/T = k$ . These two laws can be combined to form the ideal gas law, a single generalization of the behaviour of gases known as an equation of state,  $PV = nRT$ , where  $n$  is the number of gram-moles of a gas and  $R$  is called the universal gas constant. Though this law describes the behaviour of an ideal gas,



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it closely approximates  
the behaviour ...

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mass of gas is directly  
proportional to its Kelvin  
temperature if the  
pressure is kept constant.

Charles ' Law For a  
given mass of gas at  
constant temperature, the  
volume of a gas varies  
inversely with pressure

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The Ideal Gas Law relates the pressure, temperature, volume, and mass of a gas through the gas constant " R ".

$$P_{\text{total}} = P_1 + P_2 + P_3 + \dots + P_n$$

~~Gas Law's Worksheet~~  
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Section 3.2 The Gas Laws

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(pages 75 – 81) This section discusses gas pressure and the factors that affect it. It also explains the relationships between the temperature, volume, and pressure of a gas. Reading Strategy (page 75) Identifying Cause and Effect As you read, identify the variables that affect gas pressure, and write them in the ...

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~~Chapter 3 States of Matter Section 3.2 The Gas Laws~~

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Gas Laws STUDY GUIDE Due: February 12th

Ideal Gas Law Worksheet  $PV = nRT$

Use the ideal gas law, “  $PV = nRT$  ”, and the universal gas constant  $R = 0.0821 \text{ L} \cdot \text{atm}$  to solve the

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### following Answers

problems:  $K \cdot \text{mol}$  If  
pressure is needed in kPa  
then convert by  
multiplying by 1013 kPa /  
1 atm to get Ideal Gas

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Assume that the  
temperature and the  
amount of gas are  
constant in the

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