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Stoichiometry. SECTION 1. SHORT ANSWER Answer the following questions in the space provided. 1. _____ The coefficients in a chemical equation represent the (a) masses in grams of all reactants and products. (b) relative number of moles of reactants and products. (c) number of atoms of each element in each compound in a reaction.

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CHAPTER 9 REVIEW Stoichiometry
SECTION 3 PROBLEMS Write the
answer on the line to the left. Show all
your work in the space provided. 1. 88%
The actual yield of a reaction is 22 g and
the theoretical yield is 25 g. Calculate the
percentage yield. 2. 6.0 mol of N₂ are
mixed with 12.0 mol of H

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SECTION 2 continued Date Class _____

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g/L calculate the volume of this gas. 76

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CHEMISTRY a. —. 81 g 6. A car air bag
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SECTION 3 PROBLEMS Write the

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your work in the space provided. 1. 88%

The actual yield of a reaction is 22 g and
the theoretical yield is 25 g. Calculate the

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mixed with 12.0 mol of H₂ according to

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the following equation: N

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Stoichiometry. SECTION 2. PROBLEMS

Write the answer on the line to the left.

Show all your work in the space provided.

1. The following equation represents a laboratory preparation for oxygen gas:
 $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$ How many moles of O_2 form if 3.0 mol of KClO_3 are totally consumed?
2. Given the following equation: $\text{H}_2(\text{g}) + \text{F}_2(\text{g}) \rightarrow 2\text{HF}(\text{g})$

CHAPTER 9 REVIEW

This section is an arithmetic of the equation, and you should be able to work with the entire class. 3 1 Formula Mass and Mole Concept – Stoichiometry section 12.1 in the arithmetic of equation worksheet answers chemistry, Source:opentextbc.ca There are two reasons why students struggle with this

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Chapter 12.1 stoichiometry worksheet answers

1. Write the definition of reaction stoichiometry in your own words.

Introduction to Stoichiometry SECTION 9.1 amount of given substance (mol) convert into amount of unknown substance (mol) Ratios of substances in chemical reactions can be used as conversion factors. Reaction stoichiometry problems can be approached by looking

SECTION 9.1 Introduction to Stoichiometry

CHAPTER 9 REVIEW Stoichiometry SECTION 3 PROBLEMS Write the answer on the line to the left Show all your work in the space provided 1 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g Calculate the

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percentage yield 2 60 mol of N₂ are
mixed with 120 mol of H

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Answer Key

368 Chapter 11 • Stoichiometry Section
11.1.1 Objectives Describe the types of
relationships indicated by a balanced
chemical equation. State the mole ratios
from a balanced chemical equation.

Review Vocabulary reactant: the starting
substance in a chemical reaction New
Vocabulary stoichiometry mole ratio
Defining Stoichiometry

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table) Day 12 - 11/20 Worksheet:
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identify reaction types, write equations,

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balance . Day 13 - 11/26 FC –

Stoichiometry Introduction Worksheet:

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SECTION 3 PROBLEMS Write the

answer on the line to the left. Show all

your work in the space provided. 1. _____

The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol

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