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KEY. Write the empirical formula for the following compounds.

1) C_6H_6
2) CH_4
3) C_8H_{18}
4) C_4H_9
5) WO_2
6) WO_2
7) $C_2H_6O_2$
8) CH_3O
9) $X_{39}Y_{13}$
10) X_3Y_6

A compound with an empirical formula of C_2H_4O and a molar mass of 88 grams per mole. What is the molecular formula of

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this compound? C. 4. O.

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then another formula

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the formula in the

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worksheet will be a combination of numbers that will be different than the numbers on the top part of the worksheet.

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 $x(\text{empirical formula}) =$
molecular formula. also.
 $x(\text{empirical formula}$
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formula mass. 1. To determine the molecular formula of a compound, you must know the compound's molar mass. 2. Divide the molecular mass by the empirical formula mass to determine the whole number multiple (x).

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Formula 1) What is the

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compound that contains

0.783g of Carbon,

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0.196g of Hydrogen and

0.52 lg of Oxygen?

1500 q 2) What is

empirical formula of a

compound which

consists of 89.14% Au

and 10.80% of O? o,qs

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Chemical formulas can
have three components.

a) The _____ which are
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represented by an uppercase letter followed by a lowercase letter. This tells you the types of elements in the compound. b) The _____ which are numbers that are found on the lower _____-hand side

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down the chemical formula of reactants that are listed on the left side of the chemical equation. After this, you can list down the products on the right hand side of the chemical equation. There is an arrow between the sides, signaling the direction the reaction is happening in.

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their structure in space,

although there is A

formula used to indicate

not just the range of

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