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of Formation  
Reaction \u0026  
Heat of  
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|| Part - 11 ||

Chemistry ||

Class-11 ~~15.1~~



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Molar Enthalpy  
Calculations~~

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reaction  
enthalpy from  
formation  
enthalpy data.  
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Dec 19/Altered  
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Journal/Buttons  
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~~Enthalpy~~

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## **STANDARD ENTHALPY OF FORMATION List Of Enthalpies Solution**

Table 9.5.1  
Enthalpies of  
Solution at 25°C  
of Selected  
Ionic Compounds  
in Water (in  
kJ/mol) Anion.  
Cation.

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Fluoride.  
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Bromide. Iodide.  
Hydroxide.  
lithium. Chapter  
9.5: Enthalpies  
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the enthalpies  
of formation for  
the common  
cations of some  
metallic

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representative  
ketones,  
phenols,  
alcohols, and  
ethers have been  
determined in ?,  
?,?-trifluorotol  
uene, benzene,  
toluene, and  
mesitylene, and

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**Solution -**

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Solution

Enthalpies of solution may be either positive or negative - in other words, some ionic substances dissolved endothermically (for example,

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NaCl); others dissolve exothermically (for example NaOH). An infinitely dilute solution is one where there is a sufficiently large excess of water that adding any more doesn't

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Solution

Enthalpies of  
solution may be  
either positive  
or negative - in

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other words, some ionic substances dissolved endothermically (for example, NaCl); others dissolve exothermically (for example NaOH). An infinitely dilute solution is one where

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there is a Page  
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in the  
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formation for  
the common  
cations of some  
metallic  
elements in

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aqueous solution. The formation of an aqueous cation from an element in its standard state is a fairly abstract multi-step process, but it relates directly to the oxidation-reduction reactivity of

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the element and  
to the  
solubility of  
ionic compounds.  
So, this is

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Solution -  
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University**

An infinitely  
dilute solution  
is one where  
there is a



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sufficiently

large excess of  
water that

adding any more  
doesn't cause  
any further heat  
to be absorbed  
or evolved. So,  
when 1 mole of  
sodium chloride  
crystals are  
dissolved in an  
excess of water,  
the enthalpy

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change of  
solution is  
found to be  $+3.9$   
 $\text{kJ mol}^{-1}$ . The  
change is  
slightly  
endothermic, and  
so the  
temperature of  
the solution  
will be slightly  
lower than that  
of the original  
water.

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## **ENTHALPIES OF SOLUTION AND HYDRATION -**

**chemguide**

Standard  
Enthalpies of  
Formation. The  
magnitude of  $\Delta H$   
for a reaction  
depends on the  
physical states  
of the reactants  
and the products

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(gas, liquid, solid, or solution), the pressure of any gases present, and the temperature at which the reaction is carried out. To avoid confusion caused by differences in reaction

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conditions and  
ensure  
uniformity of  
data, the  
scientific  
community has  
selected ...

**5.7: Enthalpies  
of Formation -  
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Enthalpies of  
Solution 1.

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Authors: B. D.

Lamp, T.

Humphry, V. M.

Pultz and J. M.

McCormick\* Last

Update: November

13, 2013.

Introduction.

Thermochemistry

Background.

Thermochemistry

investigates the

relationship

between chemical

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**Enthalpies** and  
energy changes  
involving heat.  
It was born out  
of the practical  
problem of  
cannon making  
and today ...

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**Solution | Chem**  
**Lab**

Estimating  
enthalpies of

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Solution from  
lattice  
enthalpies and  
hydration  
enthalpies. The  
hydration  
enthalpies for  
calcium and  
chloride ions  
are given by the  
equations: The  
following cycle  
is for calcium  
chloride, and



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includes a  
lattice  
dissociation  
enthalpy of  
 $+2258 \text{ kJ mol}^{-1}$ .  
We have to use  
double the  
hydration  
enthalpy of the  
chloride ion ...

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provide these  
unique  
information  
services, Doody  
Enterprises has  
forged  
successful

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?  $t H =$  Sum of  
enthalpies of  
the product -

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Sum of the  
enthalpies of  
the reactants.

This is a brief  
outline of the  
concept of  
enthalpy.

However, our  
main aim in this  
chapter is to  
earn about the  
enthalpies of  
different  
reaction. Let us

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now look at  
these. . . .

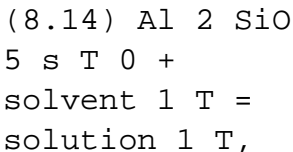
Enthalpy of  
Solution. We may  
define it as,  
"The quantity of  
heat evolved or  
. . . .

**Enthalpies for  
Different Types  
of Reactions:  
Enthalpy . . .**

For example, the

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Enthalpies of formation of  $\text{Al}_2\text{SiO}_5$  [115] can be obtained from three enthalpy of solution measurements corresponding to the following reactions:



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(8.15) Al<sub>2</sub>O<sub>3</sub>  
s T<sub>0</sub> + solvent  
1 T = solution 1  
T, (8.16) SiO<sub>2</sub>  
s T<sub>0</sub> + solvent  
1 T = solution 1  
T.

**Enthalpy of  
Solution - an  
overview |  
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The following

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Enthalpies can be used to calculate the standard enthalpy of reaction:  $\Delta H_{\text{rxn}} = \sum \Delta H_f^{\circ} \{\text{products}\} - \sum \Delta H_f^{\circ} \{\text{reactants}\}$ . The enthalpy of



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reaction is  
calculated under  
standard  
conditions  
(STP).

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**Enthalpy of**  
**Formation and**  
**Reaction |**  
**Boundless ...**  
Standard  
Enthalpies of  
Formation &

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Entropies of  
Common Compounds

Substance State

$\Delta H_f^\circ$   $S^\circ$  (kJ mol<sup>-1</sup>)

(J mol<sup>-1</sup>·K) Ag s 0

42.6 Ag<sup>+</sup> aq

105.79 72.7 AgCl

s  $\Delta H_f^\circ$  127.01 96.2

AgBr s  $\Delta H_f^\circ$  100.4

107.1 AgNO<sub>3</sub> s

$\Delta H_f^\circ$  124.4 140.9 Al

s 0 28.3 Al<sup>3+</sup> aq

$\Delta H_f^\circ$  538.4  $\Delta H_f^\circ$  321.7

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AlCl<sub>3</sub> s ?704  
110.7 Al<sub>2</sub>O<sub>3</sub> s  
?1675.7 50.9 Ba  
s 0 62.8 BaCl<sub>2</sub>  
s ?858.6 123.7  
BaCO<sub>3</sub> ...

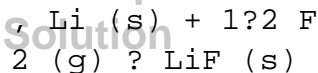
**Standard  
Enthalpies of  
Formation &  
Standard  
Entropies of ...**

For example, the  
formation of

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lithium fluoride



may be

considered as

the sum of

several steps,

each with its

own enthalpy (or  
energy,

approximately):

The standard

enthalpy of

atomization (or

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sublimation) of  
solid lithium.

The first  
ionization  
energy of  
gaseous lithium.

**Standard  
enthalpy of  
formation -  
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The heat of  
solution, like  
all enthalpy

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Enthalpies is expressed in kJ/mol for a reaction taking place at standard conditions (298.15 K and 1 bar). The heat of solution can be regarded as the sum of the enthalpy changes of three

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intermediate  
steps: The value  
of the overall  
heat of  
solution,  
 $\Delta H^{\circ}_{\text{sol}}$ ,  
is the sum of  
these individual  
steps.

**Definition of  
enthalpy of  
solution in**

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Many experimental data are tabulated according to the type of process. For example, extensive tables exist of enthalpies of vaporization ( $H$  for converting liquids to



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Enthalpies  
Solution

gases),  
enthalpies of  
fusion ( $H$  for  
melting solids),  
enthalpies of  
combustion ( $H$   
for combusting a  
substance in  
oxygen), and so  
forth.

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