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a. Describe how you would prepare 1 L of a 1 M solution of sodium chloride. The gram formula weight of sodium chloride is 58.44 g/mol. Answer: To make a 1 M solution of sodium chloride, dissolve 58.44 g sodium chloride in 500 mL water in a 1000-mL volumetric flask. When all the solid is dissolved and the solution is at room temperature,

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**Solutions** dilute to the mark and  
invert the flask several  
times to mix.

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Bonds. NaCl is made up of  
sodium cations and chloride  
anions, so it is ionic. C<sub>6</sub>  
H<sub>12</sub> O<sub>6</sub> is made up...

Structure and Molecular

Representation. Each oxygen  
and sulfur shown has 2 lone  
pairs. Each nitrogen shown  
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A 0.050 molal solution of  $\text{CH}_3\text{COOH}$  (acetic acid) in cyclohexane has a density of 0.90 g/mL. Calculate the molarity of the solution. Calculate the weight percent of the solution. Calculate the freezing point of the solution. Draw an energy level diagram showing the energy of the solute and solvent relative to the energy of the solution.

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Solution. Since  $\left(\mathrm{density} = \frac{\text{mass}}{\text{volume}}\right)$ , we need to divide the mass in grams by the volume in milliliters. In general: the number of units of B = the number of units of A  $\left(\times\right)$  unit conversion factor. The necessary conversion factors are given in Table 1.7.1: 1 lb = 453.59 g; 1 L = 1.0567 qt; 1 L = 1,000 mL.

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