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PROBLEM \(\backslash\PageIndex{3}\)

Determine the molarity for each of the following solutions: 0.444 mol of CoCl_2 in 0.654 L of solution; 98.0 g of phosphoric acid, H_3PO_4 , in 1.00 L of solution; 0.2074 g of calcium hydroxide, $\text{Ca}(\text{OH})_2$, in 40.00 mL

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of solution 10.5 kg of $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ in 18.60 L of solution;
 7.0×10^{-3} mol of I_2 in 100.0 mL of solution; 1.8×10^4 mg of HCl in 0.075 L of ...

6.1.1: Practice Problems- Solution

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PROBLEM $\backslash(\backslash$ PageIndex{3}\)

Solutions of hydrogen in palladium may be formed by exposing Pd metal to H_2 gas. The concentration of hydrogen in the palladium depends on the pressure of H_2 gas applied, but

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in a more complex fashion than can be described by Henry's law. Under certain conditions, 0.94 g of hydrogen gas is dissolved in 215 g of palladium metal.

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1. A 0.750 L aqueous solution

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contains 90.0 g of ethanol, C_2H_5OH . Calculate the molar concentration of the solution in $mol \cdot L^{-1}$.: Solution:

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