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Stoichiometry SECTION 3

PROBLEMS Write the answer on the line to the left. Show all your work in the space provided.

1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the

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percentage yield. 2.6.0 mol
of N₂ are mixed with 12.0
mol of H₂ according to the
following equation: N₂(g)
3H₂(g) ? 2NH₃(g) N

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SHORT ANSWER Answer the following questions in the space provided. 1. _____ The coefficients in a chemical equation represent the (a) masses in grams of all reactants and products. (b)

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relative number of moles of reactants and products. (c) number of atoms of each element in each compound in a reaction.

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to calculate the number of

grams, moles, or particles

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