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$[4 \times 10^{-10}] [3 \times 4 \times 10^{-10}]^3 \rightarrow 6.9 \times 10^{-34}$ Q: What are the concentrations of each of the ions in a saturated solution of PbI_2 , given that the K_{sp} of PbI_2 is 8.5×10^{-9} at 25°C . If 5g PbI_2 are dissolved in water to make 1L of solution at 25°C , would it be sat, unsat, or supersat

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Mole-mole factors come from the coefficients of a balanced chemical equation. See Section 9.2 (Page 252) . ANSWER: Correct Chapter 9 Reading Question 2 Part A Iron reacts with oxygen in the air to produce rust (Fe_2O_3) according to the reaction shown below. $4 \text{Fe} + 3 \text{O}_2 \rightarrow 2 \text{Fe}_2\text{O}_3$ How many moles of rust can be formed if 1.8 mols of iron are present? Hint 1.

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Answer Key Chapter 9 - Chemistry: Atoms First 2e | OpenStax. 1. The temperature of 1 gram of burning wood is approximately the same for both a match and a bonfire. This is an intensive property and depends on the material (wood). However, the overall amount of produced heat depends on the amount of material; this is an extensive property.

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9.1. Write the formulas for the following coordination compounds:
(i) Tetraammine diaquacobalt (III) chloride. (ii) Potassium tetracyanonickelate (II) (iii) Tris (ethane-1,2-diamine) chromium (III) chloride. (iv) Amminebromidochloridonitrito-N- platinum (II) (v) Dichloridobis (ethane-1,2-diamine) platinum (IV) nitrate.

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Chemistry (4th Edition) Burdge, Julia Publisher McGraw-Hill Publishing Company ISBN 978-0-07802-152-7

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Chapter 9 Chemical Quantities. 1. Although we define mass as the “amount of matter in a substance,” the units in which we measure mass are a human invention. Atoms and molecules react on an individual particle-by-particle basis, and we have to count individual particles when doing chemical calculations. 2.

Chapter 9 Chemical Quantities

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explore and make use of the seemingly limitless changes that matter can undergo, chemists and chemistry students often need to answer questions that

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As we noted in Chapter 9.2 the heat released by a reaction carried out at constant volume is identical to the change in internal energy (ΔE) rather than the enthalpy change (ΔH); ΔE is related to ΔH by an expression that depends on the change in the number of moles of gas during the reaction. The difference between the heat flow measured at constant volume and the enthalpy change is usually quite small, however (on the order of a few percent).

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