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single atom with a positive or negative charge resulting from the loss of gain of one or more valence

electrons and

behaves as a

unit, yet it

still has a

charge. (+ or

-). Determined

by the amount of

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ion. - single  
atom with a  
positive or  
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resulting from  
the loss or gain  
of electrons,  
respectively.

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Naming and  
Writing Formulas  
for Acids and  
Bases. An acid  
is a compound  
that produces  $H^+$   
ions when it  
dissolves in  
water. The  
formula for an

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acid normally  
starts with and  
Names And  
H. When naming  
Formulas  
acids, you  
should first  
determine  
whether or not  
the acid  
contains oxygen.

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Acids. base.

Cation. Ionic  
compounds.

Compounds that  
contain one or  
more hydrogen  
atoms and

produce... An  
ionic compound  
that produces

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hydroxide ions  
when dissolved...  
Any atom or  
group of atoms  
that has a  
positive charge.  
Compounds  
composed of  
cations and  
anions.

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SECTION 9.3

NAMING AND

WRITING FORMULAS

FOR MOLECULAR

COMPOUNDS (pages  
268–270) This

section explains  
the rules for  
naming and

# Download Ebook Chapter 9 Chemical Formulas Names And Formulas writing formulas for binary molecular compounds.

Naming Binary  
Molecular  
Compounds (pages  
268–269) 1.  
Circle the  
letter of the  
type(s) of  
elements that  
form binary  
molecular



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compounds.

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COMPOUNDS I.

Write the

formulas for

these binary

ionic compounds.

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c. potassium iodide  
2. write the formulas for the compounds formed  
c. b. f.

3. Name the following binary ionic compounds.  
sodium sulfide

g.  $\text{CuCl}$ : h.

snc. l. a.  $\text{MnO}_2$

c. d.  $\text{SrBr}_2$  221

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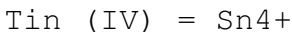
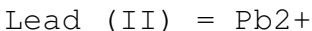
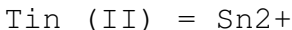
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... in  
parenthesis  
after the name  
of the metal  
(Table 9.2,  
p.255)

Predicting Ionic  
Charges Some of  
the post-  
transition

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elements also  
have more than  
one possible  
oxidation state.



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Assessment 42.

a.  $2^-$  b.  $1^+$  c.  $1^-$  d.  $3^+$  43. a.  $2^+$  b.  $2^+$  c.  $3^+$  d.  $1^+$  44. a.

barium ion b.

iodide ion c.

silver ion d.

mercury(II) ion

45. cyanide,  $\text{CN}^-$

and hydroxide,

$\text{OH}^-$  46. a.

hydroxide ion b.

lead(IV) ion c.

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sulfate ion d.  
oxide ion 47.  
zero; A compound  
is electrically  
neutral. 48. The  
symbols for the  
cation and

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## Section 9.1: Names And Formulas Naming Ions. 1.

What is the charge on the ion typically formed by each element?

a. oxygen c. sodium  
e. nickel, two electrons lost .  
b. iodine d. aluminum f.  
magnesium. 2.

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How many electrons does the neutral atom gain or lose when each ion forms? a.  $\text{Cr}^{3+}$   
c.  $\text{Li}^{+}$  e.  $\text{Cl}^{-}$  b.  
 $\text{P}^{3-}$  ...

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Tools. Copy this  
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friend; Find

other

activities; ...

In samples of

any chemical

compound, the

masses of the

elements are

always in the

same proportions

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any atom or  
group of atoms  
that has a  
negative charge  
... place cation  
name first  
followed by the  
anion name:  
naming  
polyatomic ...

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