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~~CHEMISTRY Chapter 19 Acids Bases~~
Answers

Answers Chapter 19 - Acids, Bases,
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684: 52 Answer The Arrhenius
definition states that acids are
compounds that contain hydrogen,
and when placed in aqueous solution ,
ionize to form hydrogen (H⁺

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An acid is a substance that contains hydrogen and ionizes to produce hydrogen ions in aqueous solution; A base is a substance that contains a hydroxide group and dissociates to produce a hydroxide ion in aqueous solution. Bronsted-Lowry model. A model of acids and bases in which an acid is a hydrogen-ion donor and a base is a hydrogen-ion acceptor.

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one a hydrated salt is copper sulphate crystals $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$.

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Answer The Arrhenius definition

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states that acids are compounds that contain hydrogen, and when placed in aqueous solution, ionize to form hydrogen (H^+ ions).

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two substances that are related by the loss or gain of a single hydrogen ion; ammonia (NH_3) and the ammonium ion (NH_4^+) are a conjugate acid-base pair (19.1) monoprotic acid. any acid that contains one ionizable proton (hydrogen ion); nitric acid (HNO_3) is a monoprotic acid (19.1) diprotic acid.

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balanced formula equation for the reaction that occurs between each of the following pairs of reactants. a. magnesium and nitric acid b. aluminum and sulfuric acid c. calcium carbonate and hydrobromic acid d. potassium hydrogen carbonate and hydrochloric acid

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Acid and Base Worksheet - Answers.

1) Using your knowledge of the Brønsted-Lowry theory of acids and bases, write equations for the following acid-base reactions and indicate each conjugate acid-base pair: a) $\text{HNO}_3 + \text{OH}^- \rightarrow \text{H}_2\text{O} + \text{NO}_3^-$. HNO_3 and NO_3^- make one pair OH^- and H_2O make the other.

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