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?chemical equation A

representation of a chemical

reaction with reactants on

the left, products on the

right, and an arrow

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Chapter 11 Chemical
Reactions 113 SECTION 11.1

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DESCRIBING CHEMICAL REACTIONS (pages 321-329)

This section explains how to write equations describing chemical reactions using appropriate symbols. It also describes how to write balanced chemical equations

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when given the names or formulas of the reactants and products in a chemical reaction. Writing Chemical Equations (pages 321–323) 1.

SECTION 11.1 DESCRIBING CHEMICAL REACTIONS (pages

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separating the two skeleton
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Answer Key

Section 11.1 Assessment.

Describe the steps in writing a balanced chemical equation. Write the skeleton equation for the following reactions: Heating solid copper(II) sulfide in the

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presence of oxygen gas
produces pure copper and
sulfur dioxide gas. Iron
metal and chlorine gas react
to form solid
iron(III)chloride. $CuS (s) +$
 $O_2 (g) \rightarrow Cu (s) + SO_2 (g)$ D

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Chapter 11: Chemical Reactions

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Chemical Reactions 119 GUIDED
PRACTICE PROBLEM 2 (page
324) 2. Sulfur burns in
oxygen to form sulfur
dioxide. Write a skeleton

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equation for this chemical reaction. Include appropriate symbols from table 11.1. Analyze Step 1. Identify the relevant concepts. Write the formula for each reactant and ...

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AnswersChemical Reactions.

In a chemical reaction, the reactants are written on the

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left and the products on the right. The arrow that separates them is called yield.

[PDF] Chapter 11 Chemical Reactions

CHEMICAL REACTIONS. Chapter

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11 Chemical Reactions 113.

SECTION 11.1 DESCRIBING

CHEMICAL REACTIONS (pages

321–329). This section

explains GUIDED

PRACTICE PROBLEM 17 (page

334). 17. Complete the

equations for these single-

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replacement reactions in aqueous solution. Balance each equation. Write “no reaction” if a ...

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Shapkiore

The initial step in solving a problem of this type must be to write the balanced chemical equation for the reaction. Inspection of Equation 11.4.1 shows that

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it is balanced as written,
so we can proceed to the
strategy outlined in Figure
11.4.1, adapting it as
follows:

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