

Calculating Ph Packet Answers Pogil Answer

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Calculating Ph Packet Answers Pogil Answer The pH of the solution will be governed by the following base hydrolysis equilibrium, which is just the reverse of the neutralization reaction: $\text{A}^- + \text{H}_2\text{O} \rightleftharpoons \text{HA} + \text{OH}^-$ K_b for A^- , which may be calculated from K_a of HA is $K_b = K_w / K_a = 1.00 \times 10^{-14} / 1.00 \times 10^{-5} = 1.00 \times 10^{-9}$ Making the usual

Calculating Ph Pogil Answer Key

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Keep in mind that, x is equal to the concentration of hydroxide ions now in the solution. $[\text{OH}^-] = 6 \times 10^{-4} \text{M}$. Use this value in the equation for pOH : $\text{pOH} = -\log [\text{OH}^-] = -\log (6 \times 10^{-4}) = 3.2$. Remember that the sum of pH and pOH is always 14.

Calculating pH and pOH - High School Chemistry

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