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Spectator ions are present
in acidic and basic
solutions, but do not
participate in the
neutralization reaction
between the $H^+(aq)$
(hydrogen ions) and

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OH^- (aq)(hydroxide ions). Spectator ions can be positive or negative, and they are present in quantities needed to produce electrically neutral solutions.

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Neutralization Reactions

Given the following
information, solve the
practice problems

below. In a neutral

solution the Moles of

$H^+ = \text{Moles of } OH^-$. #

Moles = Molarity x

Volume (# Moles =

$M \cdot V$) In a neutral

solution $M_A V_A = M_B V_B$

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$M_A V_A = M_B V_B$ (where M_A = Molarity of the hydrogen ion, V_A = volume of the acidic solution, M_B = Molarity of the hydroxide ion and V_B

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Together, each pair of 2
will complete a POGIL
on neutralization

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reactions. When the POGIL is completed, each group will come back to the class and we will review. The purpose of a POGIL is to preview the information and have them learn on their own by starting very basic and building on each concept as they work through the packet.

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Neutralization Reactions

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Titration

POGIL.notebook 5

March 18, 2016 In the titration of a strong acid and a strong base the pH at equivalence = 7.00 because the only major species that remains is water

Titration

POGIL.notebook March

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18, 2016

Acids and bases react with one another to yield two products: water, and an ionic compound known as a salt. This kind of reaction is called a neutralization reaction.

10.1: Introduction to

Acids and Bases -

Chemistry LibreTexts

We can't discuss acids and bases without

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Answers

talking about pH. pH
measures the
acidity/basicity of a
solution.

Introduction To Acids
And Bases Pogil
Answers

Information (pH at the
Equivalence Point) In a
strong acid - strong base
titration, neutralization
produces water and an
aqueous solution of a

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salt, whose cation and anion come from the base and acid, respectively. Neither ion is acidic or basic, so the pH is that of neutral water; i.e., 7.00.

Chem 116 POGIL

Week 11

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Neutralization Reactions

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limits book, fiction,
history, novel, scientific
research, as skillfully as
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...

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Neutralization Reactions

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For each acid-base

reaction in Model 2,

describe the role of the

Brønsted-Lowry acid in

the ion (proton) transfer

that occurs. Q cid 1-4 ±

For each acid-base

reaction in Model 2,

describe the role of the

Brønsted-Lowry base in

the pro- ton (H' ion)

transfer that occurs.

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High School Chemistry
Reactions Pogli
a.

Answers

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Give the name and the
formula of the ionic
compound produced by
neutralization reactions
between the following
acids and bases: Acid
and Base reactants nitric
acid and sodium

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hydroxide hydroiodic
acid and calcium
hydroxide magnesium
hydroxide and
hydrosulfuric acid
ammonium hydroxide
and hydrofluoric acid
barium hydroxide z and
sulfuric acid

Mrs. Zuberbuehler -
Mrs. Zuberbuehler
strong acid or any
strong base that is

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added, allowing the solution to keep a fairly constant pH. 7. Which beaker in Model 1 contains a buffer? D For this buffer

a) What species is the weak acid? HNO_2 What species is the weak base? NO_2^-

b) Write the neutralization reaction that would take place if 1.0 M NaOH was added to this buffer.

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Neutralization

POGIL Activities for
AP* Chemistry-

modified Name Buffers

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TITRATIONS In this activity we will explore titration, pH curves and acid-base indicators. We will examine two types of titrations: strong acid-strong base titration (relatively simple) and strong acid-weak

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base/weak acid-strong
base titrations (a lot
more involved).

Answers

Scarsdale Public
Schools / Overview
Lesson 1: Introduction
to Reaction Rate. Read
Chapter 17, pages 528 -
531 in the Glencoe -
Chemistry: Matter &
Change textbook. Read
Chapter 12, pages 526 -
532 (Section 12.1) in the

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Zumdahl - Chemistry
textbook. Complete the
"Unit #3: Chemical
Kinetics" notes, and
"Practice Questions".

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