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chloride, SO_2Cl_2 , is a highly reactive gaseous compound. When heated, it decomposes as follows. $\text{SO}_2\text{Cl}_2(\text{g}) \rightleftharpoons \text{SO}_2(\text{g}) + \text{Cl}_2(\text{g})$
This decomposition is endothermic. A sample of 3.509 grams of SO_2Cl_2 is placed in an evacuated 1.00-liter bulb and the temperature is raised to 375

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K. (a) What would be the ...

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vs $1/T$ $E_{act} = -R$ (slope) or $-2,303 R$
(slope) For partial credit, if the 2-point
equation is given for the activation
energy, the student may receive a
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point if it is stated that k is plotted vs $1/T$ or if $\ln K$ or $\log k$ is plotted vs T . b) five points

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A solution containing 3.23 grams of an unknown compound dissolved in 100.0 grams of water freezes at -0.97°C .

The solution does not conduct electricity. Calculate the molecular

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Answer the question. (The

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Exam Content The AP Chemistry
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