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13-3 Practice Problems

the factory is 761.3 mm Hg. The pressure inside the bag is the same. What is the pressure inside the bag of potato Chips in Pa? The same bag of potato chips from Problem 5 is shipped to a town in Colorado, where the atmospheric pressure is 99.82 kPa. What is the difference (in Pa) between the pressure in the bag and the atmospheric

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2. 30 4. 5. A gas occupies a volume of 458 mL at a pressure of 1.01 kPa and temperature of 295 K. When the pressure is changed, the volume becomes 477 mL. If there has been no change in temperature, what is the new pressure? 13 3

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and temperature of 295-K. When the pressure is changed, the volume becomes 477-mL. If there has been no change in temperature, what is the new pressure? 6. What will be the volume of a gas sample at 309-K if its volume at 215-K is 3.42-L?

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13.3: Factors Affecting Solubility Conceptual Problems Use the kinetic molecular theory of gases discussed in Chapter 10 to explain why the solubility of virtually all gases in liquids decreases with increasing temperature.

13.E: Properties of Solutions (Exercises) - Chemistry ...
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This is a collection of worked general chemistry and introductory chemistry problems, listed in alphabetical order. Included are printable pdf chemistry worksheets so you can practice problems and then check your answers. You may also browse chemistry problems according to the type of problem.

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Practice Chemistry with Worked Chemistry Problems

13—4 Practice Problems 1. What volume would be occupied by 100. g of oxygen gas at a pressure of 1.50 atm and a of 250C? $n = \frac{m}{M} = \frac{100 \text{ g}}{32 \text{ g/mol}} = 3.125 \text{ mol}$ (IF $\cdot 273\text{k}$) 2. An air-filled balloon has a volume of 5 L at 0.94 and 250C. Soon after, the pressure changes to 0.99 and the temperature changes to 0C. What is the new volume of the balloon? $\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$ $\frac{0.94 \text{ atm} \cdot 5 \text{ L}}{298 \text{ K}} = \frac{0.99 \text{ atm} \cdot V_2}{273 \text{ K}}$ $V_2 = 1.26 \text{ L} \dots$

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- Chapter 2. Chapter 2: Measurement and Problem Solving.

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2.1: Taking Measurements.

13.8: Solution Stoichiometry - Chemistry LibreTexts

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10-3 Practice Problems (Continued) 13. Find the empirical formula of a compound, given that the compound is found to be 47.9 percent zinc and 52.1 percent chlorine by mass.
20.

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